Centre No.						Paper Reference							Surname	Initial(s)
Candidate No.						6	C	H	0	1	/	1	Signature	

Paper Reference(s)

6CH01/1

Edexcel GCE

Chemistry

Advanced Subsidiary

Unit 1: The Core Principles of Chemistry

Sample Assessment Material

Time: 1 hour 15 minutes

Examiner's use only

Team Leader's use only

Question Number	Leave Blank
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Total	

Materials required for examination
Nil

Items included with question papers
Nil

Instructions to Candidates

In the boxes above, write your centre number, candidate number, your surname, initial(s) and signature.
Check that you have the correct question paper.
Answer ALL the questions. Write your answers in the spaces provided in this question paper.
Some questions must be answered with a cross in a box (☒). If you change your mind, put a line through the box (~~☒~~) and then mark your new answer with a cross (☒).
Do not use pencil. Use black or blue ink.

Information for Candidates

The marks for individual questions and the parts of questions are shown in round brackets: e.g. **(2)**.
There are 19 questions in this question paper. The total mark for this paper is 80.
There are 28 pages in this question paper. Any blank pages are indicated.
Candidates may use a calculator.

Advice to Candidates

Quality of written communication will be taken into account in the marking of your responses to Questions 15(a), 16(d), 18(a)(iv), 18(b)(i) and 18(b)(ii). These questions are indicated with an asterisk. Quality of written communication includes clarity of expression, the structure and presentation of ideas and grammar, punctuation and spelling.

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SECTION A

Answer ALL the questions in this section. You should aim to spend no more than 25 minutes on this section. For each question, select one answer from A to D and put a cross in the box (☒). If you change your mind, put a line through the box (~~☒~~) and then mark your new answer with a cross (☒).

Use the Periodic Table as a source of data.

1. Going across a period in the Periodic Table from left to right, the **general** trend is that

☒ **A** the bonding in the element itself changes from ionic to covalent

☒ **B** the number of neutrons in the nucleus increases

☒ **C** the first ionisation energy decreases

☒ **D** the metallic character increases

Q1

(Total 1 mark)

2. The electron configurations of argon, iron, chlorine and one other element are given below, but not in order. Which one represents the unnamed element?

☒ **A** $1s^2\,2s^2\,2p^6\,3s^2\,3p^6\,3d^6\,4s^2$

☒ **B** $1s^2\,2s^2\,2p^6\,3s^2\,3p^6$

☒ **C** $1s^2\,2s^2\,2p^6\,3s^2\,3p^4$

☒ **D** $1s^2\,2s^2\,2p^6\,3s^2\,3p^5$

Q2

(Total 1 mark)

3. Buckminsterfullerene is a carbon molecule with formula C_{60} which can trap metal ions in its structure. Which of the following compounds of buckminsterfullerene would give a line of mass/charge ratio at 837.3 in a mass spectrometer?

☒ **A** Na_4C_{60}

☒ **B** K_3C_{60}

☒ **C** Ca_3C_{60}

☒ **D** AgC_{60}

Q3

(Total 1 mark)

Leave blank

4. This question is about the following equations:

A $Cu(NO_3)_2(s) \rightarrow CuO(s) + 2NO_2(g) + O_2(g)$

B $2HCl(aq) + CuO(s) \rightarrow H_2O(l) + CuCl_2(aq)$

C $C_4H_9OH(l) + 6O_2(g) \rightarrow 4CO_2(g) + 5H_2O(l)$

D $C_8H_{18}(l) + 8½O_2(g) \rightarrow 8CO(g) + 9H_2O(l)$

(a) Which equation is **not** balanced?

☒ **A**

☒ **B**

☒ **C**

☒ **D**

(1)

(b) Which equation shows incomplete combustion?

☒ **A**

☒ **B**

☒ **C**

☒ **D**

(1) **Q4**

(Total 2 marks)

Use this space for any rough working. Anything you write in this space will gain no credit.

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5. Which of the equations shown below represents the reaction for which ΔH is the standard enthalpy change of formation, $\Delta H_{f\,298}^{\ominus}$, for ethanol, C_2H_5OH. Ethanol melts at 156 K and boils at 352 K.

☒ **A** $2C(g) + 6H(g) + O(g) \rightarrow C_2H_5OH(g)$

☒ **B** $2C(s) + 3H_2(g) + O_2(g) \rightarrow C_2H_5OH(l)$

☒ **C** $2C(s) + 3H_2(g) + O(g) \rightarrow C_2H_5OH(g)$

☒ **D** $2C(s) + 3H_2(g) + \frac{1}{2}O_2(g) \rightarrow C_2H_5OH(l)$

Q5

(Total 1 mark)

6. Use the data about four fuels given below to answer this question.

Fuel	Formula	Name	Enthalpy change of combustion /kJ mol^{-1}	Molar mass /g mol^{-1}
A	CH_4	methane	–890	16
B	CH_3OH	methanol	–726	32
C	C_3H_8	propane	–2219	44
D	C_4H_{10}	butane	–2877	58

(a) Which fuel, A, B, C or D, produces most energy per gram on complete combustion?

☒ **A**

☒ **B**

☒ **C**

☒ **D**

(1)

(b) Scientists give governments advice on technical issues. What information would scientists use when advising governments on the choice of one of these fuels, if the aim was to minimise carbon dioxide production?

☒ **A** mass of carbon per gram of fuel

☒ **B** mass of carbon per kilojoules produced

☒ **C** number of kilojoules produced per gram

☒ **D** number of kilojoules produced per mole

(1) **Q6**

(Total 2 marks)

Leave blank

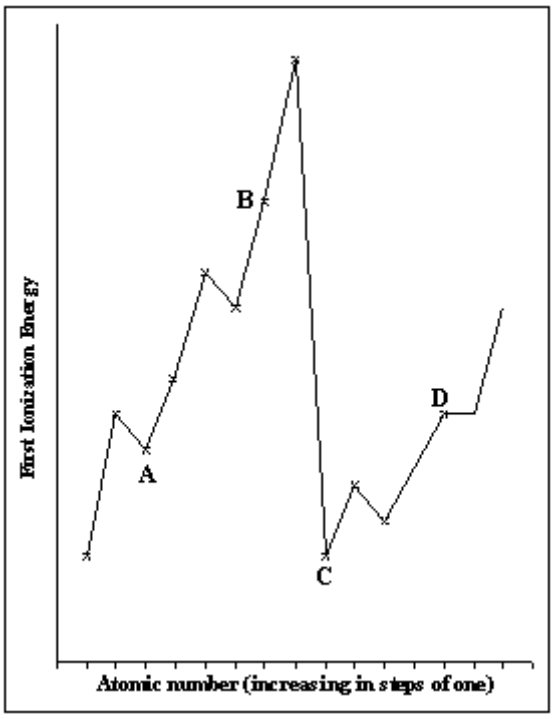
7. Which of the following equations represents the first ionisation of sulfur?

- ☒ **A** $S(s) + e^- \rightarrow S^-(g)$
- ☒ **B** $S(g) + e^- \rightarrow S^-(g)$
- ☒ **C** $S(s) \rightarrow S^+(g) + e^-$
- ☒ **D** $S(g) \rightarrow S^+(g) + e^-$

Q7

(Total 1 mark)

8. Which element marked on this graph is a halogen?



- ☒ **A**
- ☒ **B**
- ☒ **C**
- ☒ **D**

Q8

(Total 1 mark)

Leave blank

9. Question 9 is about the following ionisation energy sequences.

The values are all in kJ mol^{-1}.

A	1400	1000	950	830	700
B	420	3100	4400	5900	8000
C	1000	1250	1520	420	590
D	1520	2700	3900	5800	7200

Select from A to D the sequence which is most likely to represent the following:

(a) The first ionisation energies of five consecutive members of the same group in the Periodic Table, in order of increasing atomic number.

☒ **A**

☒ **B**

☒ **C**

☒ **D**

(1)

(b) The first five ionisation energies of an s-block element.

☒ **A**

☒ **B**

☒ **C**

☒ **D**

(1)

(c) The first five ionisation energies of a noble gas.

☒ **A**

☒ **B**

☒ **C**

☒ **D**

(1) **Q9**

(Total 3 marks)

Leave blank

10. Question 10 is about four hydrocarbons with molecular formulae as shown.

A C_2H_2

B C_3H_6

C C_3H_8

D C_4H_{10}

(a) Which hydrocarbon has the same empirical formula as its molecular formula?

☒ **A**

☒ **B**

☒ **C**

☒ **D**

(1)

Use this space for any rough working. Anything you write in this space will gain no credit.

(b) Which has a molecular ion in the mass spectrum at mass/charge ratio = 58?

☒ **A**

☒ **B**

☒ **C**

☒ **D**

(1)

Leave blank

(c) Which is neither an alkane nor an alkene?

☒ **A**

☒ **B**

☒ **C**

☒ **D**

(1)

(d) Which could be 2-methylpropane?

☒ **A**

☒ **B**

☒ **C**

☒ **D**

(1) **Q10**

(Total 4 marks)

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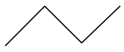
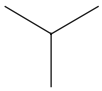
11. Question 11 is about the following organic compounds with skeletal formulae as shown:

A 

B 

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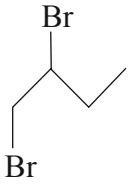
D 

(a) Which compound could be made from one of the others in an addition reaction?

☒ **A**

☒ **B**

☒ **C**

☒ **D**

(1)

Use this space for any rough working. Anything you write in this space will gain no credit.

Leave blank

(b) Which compound has E–Z isomers?

☒ **A**

☒ **B**

☒ **C**

☒ **D**

(1) **Q11**

(Total 2 marks)

12. Chemists investigating the mechanism of the reaction of ethene and bromine thought that the first step was the addition of Br^+. To test this, they reacted bromine with ethene in the presence of sodium chloride.

If their theory about the first step of the reaction was correct, which product might form as well as 1,2-dibromoethane?

☒ **A** CH_2BrCH_2Na

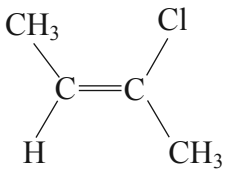
☒ **B** CH_2BrCH_2Cl

☒ **C** CH_2ClCH_2Cl

☒ **D** CH_2NaCH_2Na **Q12**

(Total 1 mark)

13. Which of the following is the correct name for the compound below?



☒ **A** Z-3-chlorobut-2-ene

☒ **B** E-3-chlorobut-2-ene

☒ **C** E-2-chlorobut-2-ene

☒ **D** Z-2-chlorobut-2-ene **Q13**

(Total 1 mark)

TOTAL FOR SECTION A: 21 MARKS

Leave blank

SECTION B

Answer ALL the questions. Write your answers in the spaces provided.

14. Copper(II) sulfate solution can be prepared from solid copper(II) carbonate by reaction with hot dilute sulfuric acid.

(a) Write the equation for the reaction, including state symbols.

(1)

(b) The experiment was carried out using 0.025 moles of sulfuric acid of concentration 2.0 mol dm^{-3}. What volume of this sulfuric acid was used?

(1)

(c) (i) It is usual to react the sulfuric acid with a slight excess of copper(II) carbonate.

Calculate the mass of copper(II) carbonate needed if a 10% excess is required.

[Molar mass of copper(II) carbonate = 123.5 g mol^{-1}]

(2)

Leave blank

(ii) A student doing this experiment chose to use a balance reading to 0.01 g in an attempt to work accurately.

Was this choice of balance necessary from the point of view of accuracy? Explain your answer.

.....

.....

(1)

(d) The sulfuric acid is heated to boiling and the copper(II) carbonate is added in small portions.

State the next step needed to prepare pure copper(II) sulfate solution. Justify your answer.

.....

.....

.....

(1)

(e) When the solution of copper(II) sulfate is allowed to crystallise, the crystals which are produced have the formula $CuSO_4.5H_2O$.

(i) What is the molar mass of $CuSO_4.5H_2O$?

(1)

(ii) 3.98 g of $CuSO_4.5H_2O$ crystals were obtained. Calculate the percentage yield in this experiment.

(2) **Q14**

(Total 9 marks)

Leave blank

15. This question is about magnesium and magnesium oxide.

*(a) Describe the bonding in magnesium and explain why it is a good conductor of electricity.

.....

.....

.....

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.....

.....

(3)

(b) Draw a diagram (using dots or crosses) for the ions in magnesium fluoride showing **all** the electrons and the ionic charges on:

(i) the magnesium ion

(1)

(ii) the fluoride ion.

(1)

(c) Under what conditions does magnesium fluoride conduct electricity?

Explain your answer.

.....

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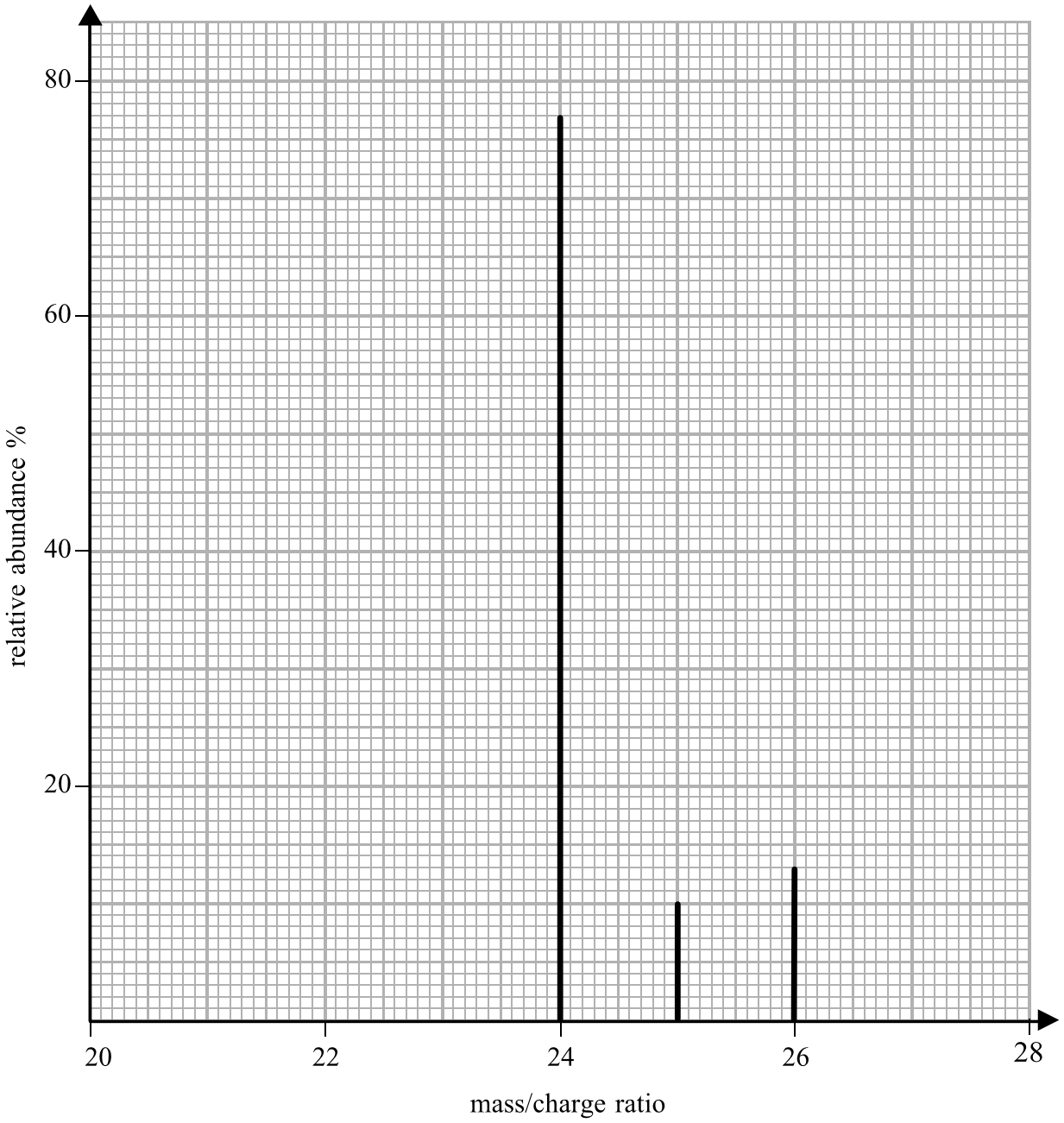
(1)

Leave blank

(d) The mass spectrum of a sample of magnesium is shown below.



(i) Use the data above to estimate the percentage isotopic composition of the sample of magnesium. Hence calculate the average atomic mass of the sample of magnesium.

(2)

Leave blank

(ii) Why do the three isotopes have the same chemical properties?

.....

.....

(1)

(e) (i) Oceanographers studying plankton found that a sample of seawater contained 1.20 nanomol dm^{-3} of chlorophyll, $C_{55}H_{77}MgN_4O_5$. (1 nanomol = 1×10^{-9} mol)

What mass of magnesium would be present in 1.00 cm^3 of this sample of seawater? Give your answer to **three** significant figures.

(2)

(ii) X-ray diffraction can be used to locate atoms or ions in molecules like chlorophyll. X-rays are scattered by the electrons in atoms and ions. In chlorophyll the atoms of one of the elements still cannot be located with certainty by this technique.

Suggest which element is most difficult to locate.

.....

.....

(1)

Q15

(Total 12 marks)

Leave blank

16. Airbags, used as safety features in cars, contain sodium azide, NaN_3. An airbag requires a large volume of gas to be produced in a few milliseconds. The gas is produced in this reaction:

$$2NaN_3(s) \rightarrow 2Na(s) + 3N_2(g) \qquad \Delta H \text{ is positive}$$

When the airbag is fully inflated, 50 dm^3 of nitrogen gas is produced.

(a) Calculate the number of molecules in 50 dm^3 of nitrogen gas under these conditions.

[The Avogadro constant = 6.02×10^{23} mol^{-1}. The molar volume of nitrogen gas under the conditions in the airbag is 24 dm^3 mol^{-1}].

(2)

(b) Calculate the mass of sodium azide, NaN_3, that would produce 50 dm^3 of nitrogen gas.

(3)

Leave blank

(c) What will happen to the temperature in the airbag when the reaction occurs?

.....

.....

(1)

*(d) The airbag must be strong enough not to burst in an accident. An airbag which has burst in an accident is hazardous if the sodium azide in it has decomposed.

Explain why this is so.

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(2) **Q16**

(Total 8 marks)

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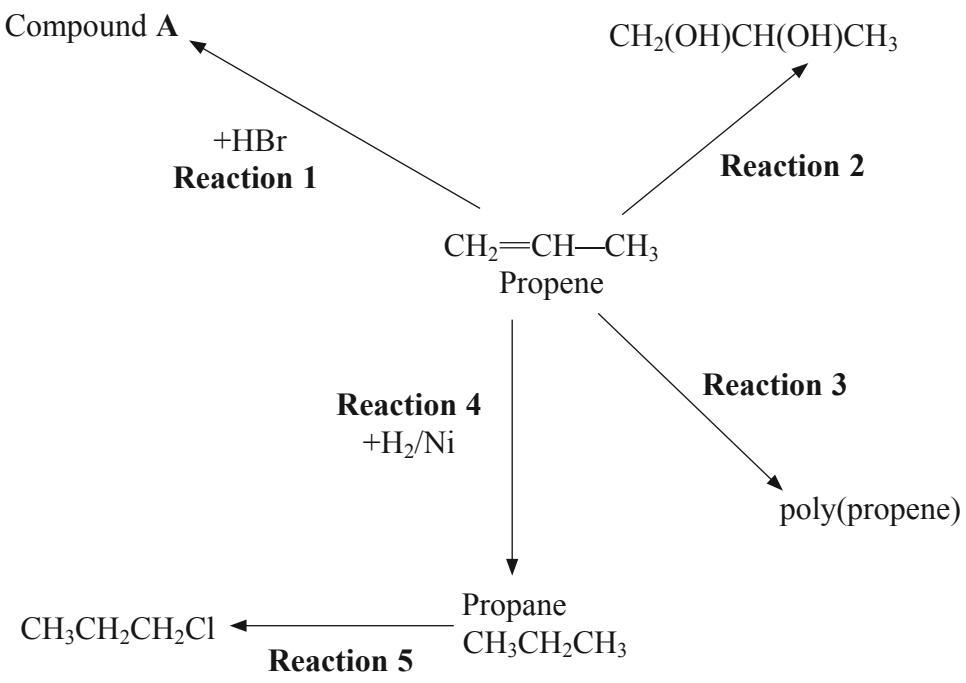
17. Propene can be used to make other important chemical products. The processes involved can be summarised in the diagram:



(a) (i) Give the mechanism for **Reaction 1.**

(3)

Leave blank

(ii) Explain why compound **A** and **not** its structural isomer is the major product in **Reaction 1**.

.....

.....

.....

(1)

(iii) Name compound **A** formed in **Reaction 1**.

Name .....

(1)

(b) What is added in **Reaction 2** to make the product $CH_2(OH)CH(OH)CH_3$?

.....

(1)

(c) Complete the balanced equation for the formation of poly(propene) in **Reaction 3** using **displayed** formulae.

$n(CH_2{=}CHCH_3) \rightarrow$

(2)

(d) Poly(propene) fibres can be used to make fleece which is used at several horse racing courses to prevent the ground becoming frozen.

State **one** advantage of using poly(propene) instead of natural fibres of similar cost.

.....

.....

(1)

Leave blank

(e) (i) One stage in the mechanism of **Reaction 5** is shown below.

$CH_3CH_2CH_3 + Cl^{\bullet} \rightarrow CH_3CH_2CH_2^{\bullet} + HCl$

What is this step?

.....

(1)

(ii) Give the name **or** formula of the trace product present in the final mixture which gives evidence for this mechanism.

.....

(1)

Q17

(Total 11 marks)

18. A student investigated a reaction which could be used to warm up coffee in self-heating cans.

$$Mg(s) + Cu(NO_3)_2(aq) \rightarrow Mg(NO_3)_2(aq) + Cu(s)$$

In the self-heating cans, the bottom has a compartment containing copper(II) nitrate solution. When a button on the bottom of the can is pressed, the magnesium powder is released into the compartment where it reacts with the copper(II) nitrate solution.

(a) A student investigated the enthalpy change for this reaction by measuring

$50.0\ cm^3$ of $0.300\ mol\ dm^{-3}$ copper(II) nitrate solution into a $100\ cm^3$ beaker and adding 1 g (an excess) of magnesium powder.

The results are shown below.

Temperature of copper(II) nitrate solution at start	= 22 °C
Temperature of mixture after reaction	= 43 °C

Leave blank

(i) Calculate the energy change which took place. The specific heat capacity of the solution is 4.20 J $g^{-1}K^{-1}$.

Which is the correct value for the energy change in joules?

(1)

(ii) How many moles of copper(II) nitrate were used in the experiment?

(1)

(iii) Calculate the enthalpy change for the reaction. You should include a sign and units in your answer.

(2)

Leave blank

*(iv) Suggest **two** changes you would make to **the equipment** used in order to improve the accuracy of the result.

.....

.....

.....

.....

(2)

(b) The ionic equation for the reaction is shown below:

$Mg(s) + Cu^{2+}(aq) \rightarrow Mg^{2+}(aq) + Cu(s)$ $\Delta H = -532$ kJ mol^{-1}

Would the following affect the value of the experimental result?

Explain your answer, stating the effect, if any, on the value of the enthalpy change obtained.

*(i) The student used 2 g rather than 1 g of magnesium.

.....

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.....

(2)

*(ii) The heat losses that occurred from the student's beaker.

.....

.....

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(2)

Leave blank

(c) The temperature in the self-heating can needs to increase by 60 °C to produce a hot drink.

Suggest a change you could make to the mixture in the experiment in (a) to produce a greater temperature rise. You are **not** expected to do a calculation.

.....

.....

(1)

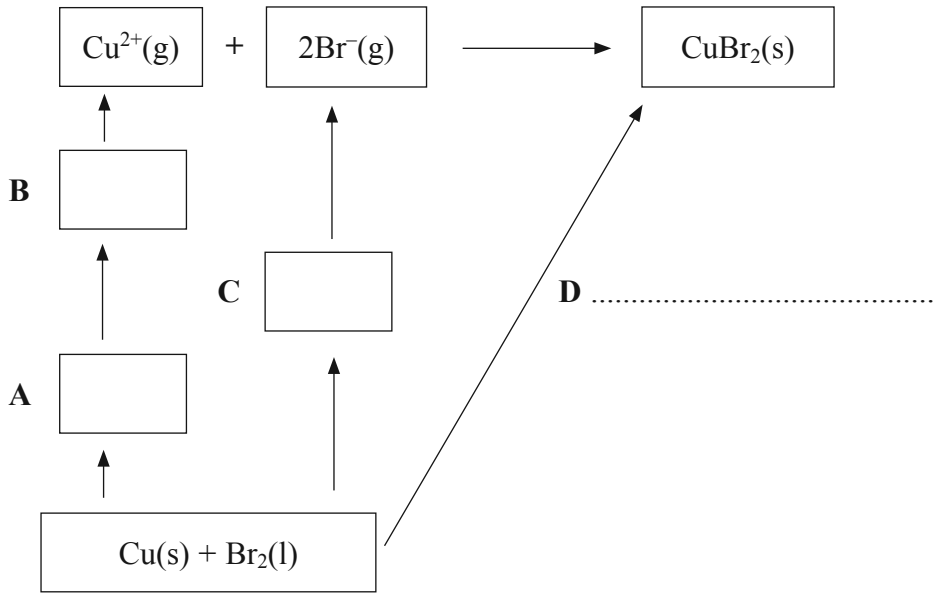
Q18

(Total 11 marks)

19. The following data can be used in a Born-Haber cycle for copper(II) bromide, $CuBr_2$.

Enthalpy change of atomisation of bromine $\Delta H^{\ominus}_{at}[\frac{1}{2}Br_{2(l)}]$	+111.9 kJ mol^{-1}
Enthalpy change of atomisation of copper, $\Delta H^{\ominus}_{at}[Cu(s)]$	+338.3 kJ mol^{-1}
First ionisation energy of copper, E_{m1} [Cu(g)]	+746.0 kJ mol^{-1}
Second ionisation energy of copper, E_{m2} [Cu(g)]	+1958.0 kJ mol^{-1}
Electron affinity of bromine, E_{aff}[Br(g)]	–342.6 kJ mol^{-1}
Enthalpy change of formation of $CuBr_2$(s), $\Delta H_f^{\ominus}$ [$CuBr_2$(s)]	–141.8 kJ mol^{-1}

(a) On the following outline of a Born-Haber cycle complete the boxes **A**, **B**, and **C** by putting in the formula and state symbol for the appropriate species and writing the name of the enthalpy change **D**.



(3)

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(b) Use the data to calculate a value for the lattice energy of copper(II) bromide.

Give a sign and units in your answer.

(3)

(c) When the lattice energy of copper(II) bromide is calculated from ionic radii and charges, the result is a value numerically about 10% less than the one obtained from the Born-Haber cycle.

(i) What does this suggest about the nature of the bonding in copper(II) bromide?

.....

.....

(1)

(ii) Draw a diagram to show how the smaller copper ion alters the shape of the larger bromide ion.

(1) **Q19**

(Total 8 marks)

TOTAL FOR SECTION B: 59 MARKS
TOTAL FOR PAPER: 80 MARKS

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The Periodic Table of Elements

1	2											3	4	5	6	7	0 (8)
							1.0 H hydrogen 1										(18) 4.0 He helium 2
(1)	(2)		Key: relative atomic mass / atomic symbol / name / atomic (proton) number									(13)	(14)	(15)	(16)	(17)	
6.9 Li lithium 3	9.0 Be beryllium 4											10.8 B boron 5	12.0 C carbon 6	14.0 N nitrogen 7	16.0 O oxygen 8	19.0 F fluorine 9	20.2 Ne neon 10
23.0 Na sodium 11	24.3 Mg magnesium 12	(3)	(4)	(5)	(6)	(7)	(8)	(9)	(10)	(11)	(12)	27.0 Al aluminium 13	28.1 Si silicon 14	31.0 P phosphorus 15	32.1 S sulfur 16	35.5 Cl chlorine 17	39.9 Ar argon 18
39.1 K potassium 19	40.1 Ca calcium 20	45.0 Sc scandium 21	47.9 Ti titanium 22	50.9 V vanadium 23	52.0 Cr chromium 24	54.9 Mn manganese 25	55.8 Fe iron 26	58.9 Co cobalt 27	58.7 Ni nickel 28	63.5 Cu copper 29	65.4 Zn zinc 30	69.7 Ga gallium 31	72.6 Ge germanium 32	74.9 As arsenic 33	79.0 Se selenium 34	79.9 Br bromine 35	83.8 Kr krypton 36
85.5 Rb rubidium 37	87.6 Sr strontium 38	88.9 Y yttrium 39	91.2 Zr zirconium 40	92.9 Nb niobium 41	95.9 Mo molybdenum 38	[98] Tc technetium 43	101.1 Ru ruthenium 44	102.9 Rh rhodium 45	106.4 Pd palladium 46	107.9 Ag silver 47	112.4 Cd cadmium 48	114.8 In indium 49	118.7 Sn tin 50	121.8 Sb antimony 51	127.6 Te tellurium 52	126.9 I iodine 53	131.3 Xe xenon 54
132.9 Cs caesium 55	137.3 Ba barium 56	138.9 La* lathanum 57	178.5 Hf hafnium 72	180.9 Ta tantalum 73	183.8 W tungsten 74	186.2 Re rhenium 75	190.2 Os osmium 76	192.2 Ir iridium 77	195.1 Pt platinum 78	197.0 Au gold 79	200.6 Hg mercury 80	204.4 Tl thallium 81	207.2 Pb lead 82	209.0 Bi bismuth 83	[209] Po polonium 84	[210] At astatine 85	[222] Rn radon 86
[223] Fr francium 87	[226] Ra radium 88	[227] Ac* actinium 89	[261] Rf rutherfordium 104	[262] Db dubnium 105	[266] Sg seaborgium 106	[264] Bh bohrium 107	[277] Hs hassium 108	[268] Mt meitnerium 109	[271] Ds darmstadtium 110	[272] Rg roentgenium 111							

Elements with atomic numbers 112-116 have been reported but not fully authenticated

* Lanthanide series	140 Ce cerium 58	141 Pr praseodymium 59	144 Nd neodymium 60	[147] Pm promethium 61	150 Sm samarium 62	152 Eu europium 63	157 Gd gadolinium 64	159 Tb terbium 65	163 Dy dysprosium 66	165 Ho holmium 67	167 Er erbium 68	169 Tm thulium 69	173 Yb ytterbium 70	175 Lu lutetium 71
* Actinide series	232 Th thorium 90	[231] Pa protactinium 91	238 U uranium 92	[237] Np neptunium 93	[242] Pu plutonium 94	[243] Am americium 95	[247] Cm curium 96	[245] Bk berkelium 97	[251] Cf californium 98	[254] Es einsteinium 99	[253] Fm fermium 100	[256] Md mendelevium 101	[254] No nobelium 102	[257] Lr lawrencium 103